

Physics 2415: General Physics II

Lecture 1

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Basic Outline

The course has three main parts, each about a month:

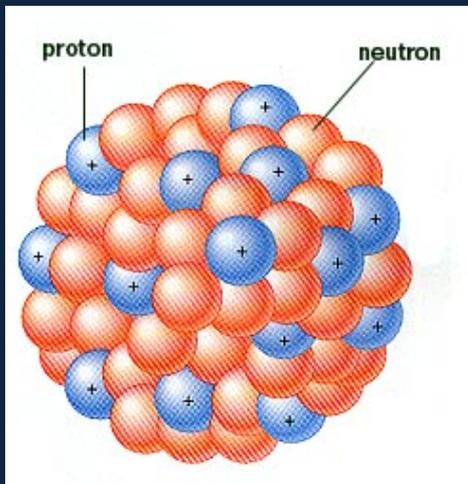
1. Electrostatics.
2. Magnetism and Magnetic Induction.
3. Electromagnetic Waves.

We'll cover some other waves, and optical phenomena, in the last few lectures. We'll also cover DC and AC circuits in the course.

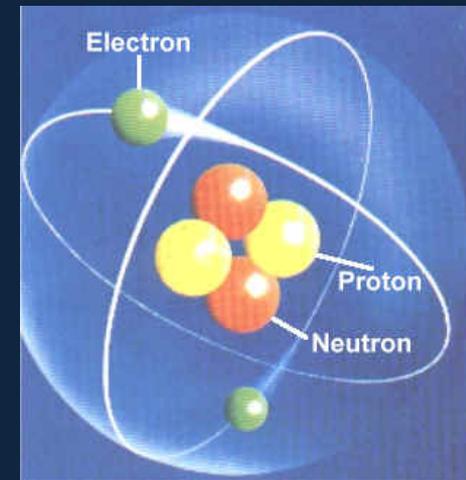
Part I: Electrostatics

1. Electric Charge and Electric Field
2. Gauss' Law
3. Electric Potential
4. Capacitance and Dielectrics

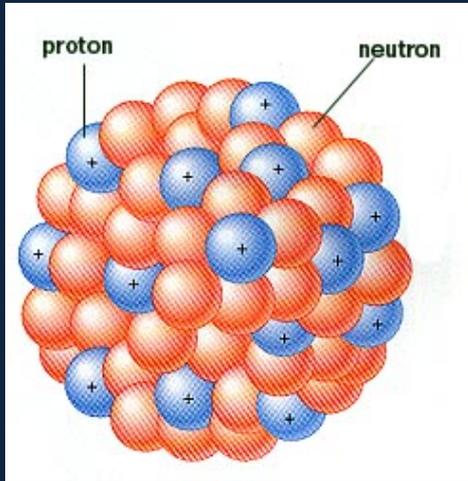
(Then we'll do DC circuits.)



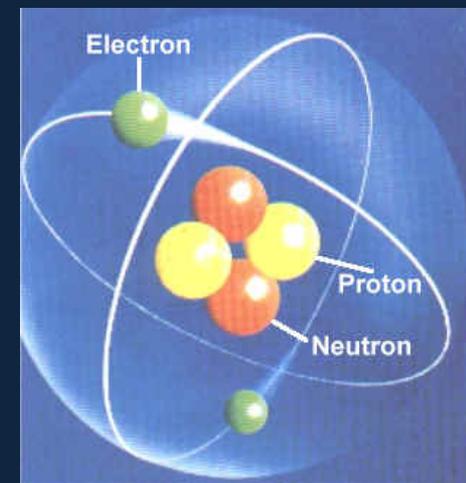
Nuclei and Atoms



- An atom consists of a positively charged nucleus surrounded by a cloud of negatively charged electrons. The total electric charge on an un-ionized atom is exactly zero.
- **Unlike charges attract:** the nucleus keeps the electrons in orbit by electrical attraction, just as the planets are held by the Sun's gravitational attraction. But...



Nuclei and Atoms



- Unlike planets around the Sun, electrons are only allowed to circle in **certain orbits**.

(Why this is true is explained by quantum mechanics, and is the basis of the Periodic Table.)

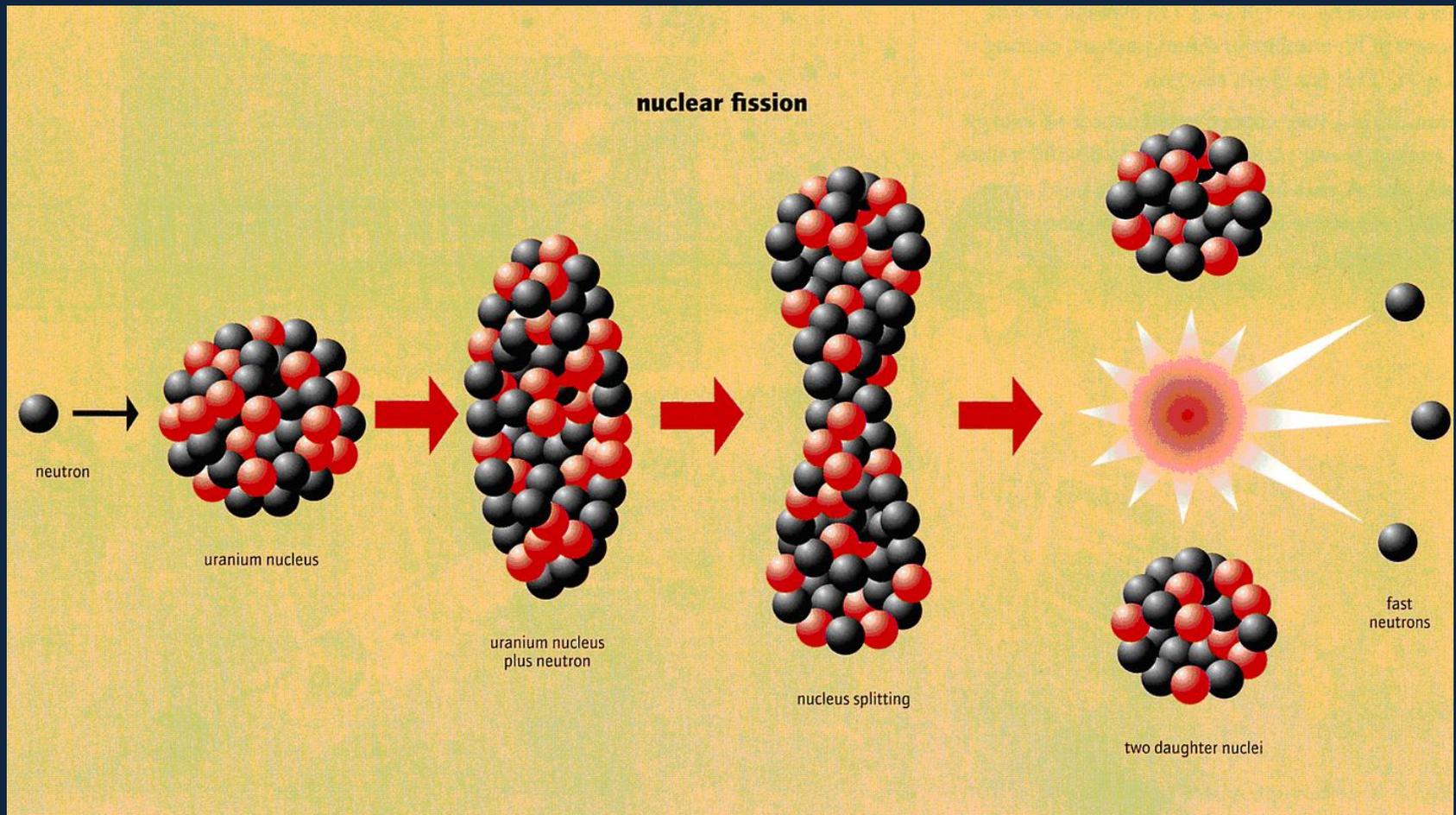
- The nucleus contains N nucleons, $N = Z + A$, Z protons each having electric charge e , A neutrons, having no charge.
- The atom contains Z electrons, charge $-e$.
- **The protons in the nucleus all repel each other!**

What Holds the Nucleus Together?

- The nucleons have a strong but **very short range** attraction, called the **nuclear force**, it's like a thin **layer of glue** around each nucleon.
- The electrostatic repulsion is **longer range**, and so is more important in bigger nuclei, unlike the nuclear “glue”, which only attracts neighbors.
- **The electrostatic repulsion is directly responsible for there being only about 100 chemical elements:** beyond that, the nuclear glue is overcome and the nucleus flies apart.

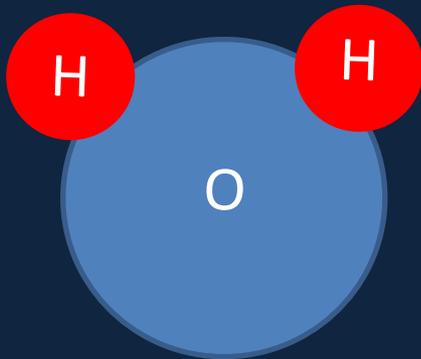
Electrostatics in Action...

A slow incoming neutron starts a Uranium nucleus wobbling like a drop of water, electrostatic repulsion drives it apart. Larger nuclei will break up without an ingoing neutron.



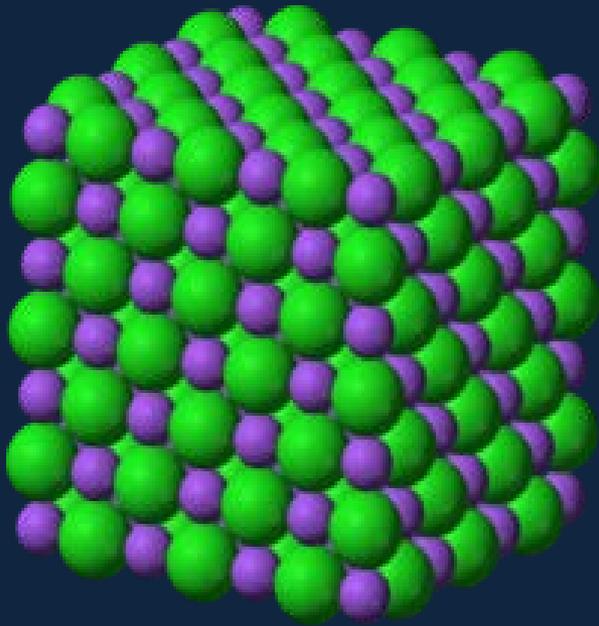
What Holds *Molecules* Together?

- Molecules are held together by **electrostatic attraction**, but the details are complicated because allowed electron distribution patterns are determined by quantum mechanics.
- Example: the Mickey Mouse water molecule:



The hydrogen atom's single electron is drawn towards the oxygen nucleus, so the H's have net positive charge, and are electrostatically attracted to the O.

What Holds *Solids* Together?



- For ordinary salt, **mainly electrostatics**: the Na atom loses an electron to become Na^+ , an ion, the Cl gains to become Cl^- , the ions form a cubic lattice where all the nearest neighbors of a Cl^- are Na^+ 's.

Contact Electrification

- If two dissimilar surfaces are placed in close contact, usually some bonding occurs between atoms, and therefore some **electron transfer takes place**—so when the surfaces are moved out of contact, one will have an excess of electrons, one a deficiency.
- **True contact** at the atomic scale is best achieved by **rubbing** the surfaces together.

Triboelectric Series

- Countless experiments in rubbing things together—cats, balloons, etc.—have established the table on the right.
- For two different surfaces, the one higher in the table will lose electrons and become positively charged.

•(Note: tribo is from the Greek for rub.)

•Electrostatic problem...

<http://www.youtube.com/watch?v=j1dTSuwz0R8>

Air	(+) P O S I T I V E
Skin (dry)	
Glass	
Human Hair	
Mica	
Nylon	
Wool	
Cat Fur	
Lead	
Silk	
Aluminum	
Paper	
Cotton	
Steel	
Wood	
Lucite	
Amber	N E G A T I V E (-)
Rubber Balloon	
Hard Rubber	
Mylar®	
Epoxy glass	
Nickel	
Copper	
Silver	
Gold, Platinum	
Polyester	
Polystyrene	
Orlon, Acrylic	
Polyester	
Cellophane Tape	
Polyurethane	
Polyethylene	
Polypropylene	
Polyimide (Kapton®)	
PVC, Vinyl	
Teflon	
Silicone Rubber	

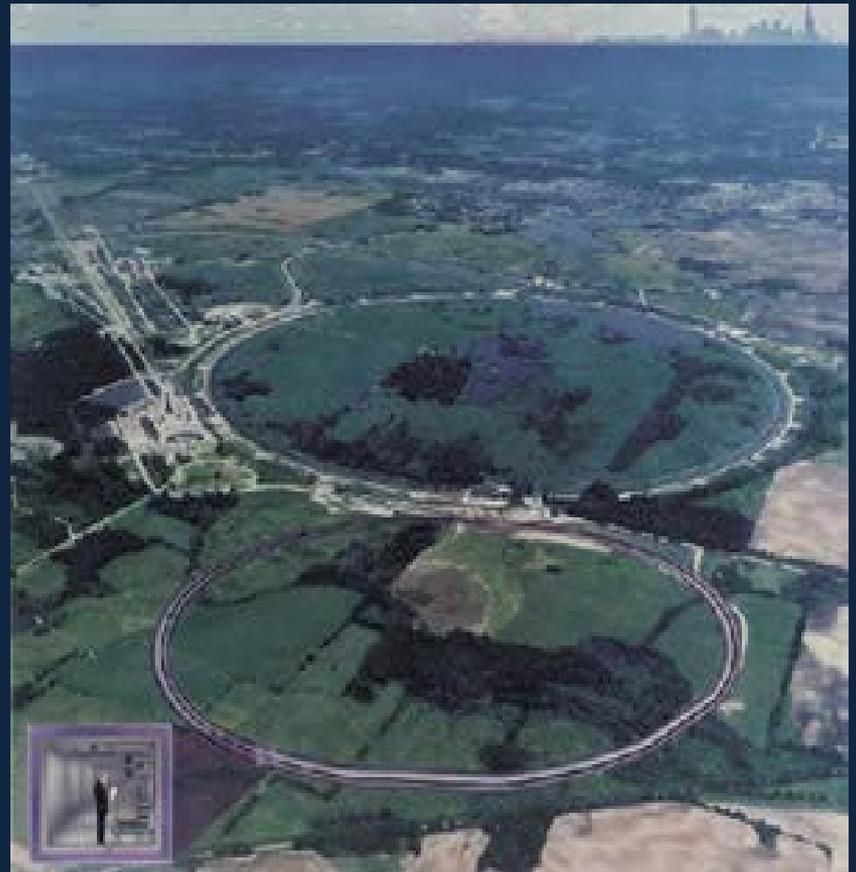
Positive and Negative Electricity

- Ben Franklin did many electrical experiments, some of them crazy, and coined the terms positively charged (meaning, we now know, deficient in electrons) and negatively charged (having extra electrons).



Electric Charge is Always Conserved

- That's obvious when we're talking about electron transfer by rubbing—but particle accelerators can *create* protons! It's found that an *antiproton*, with negative charge, is always created at the same time.
- **Creating matter means creating antimatter!**



Conductors and Insulators

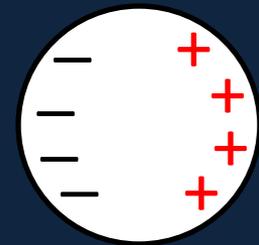
- Some solids—like aluminum—have electrons free to move from atom to atom. These are **conductors**. They're usually shiny, because the freely moving electrons reflect light very well.
- In solids like NaCl, or glass, etc., the electrons are tied to atoms, no charge can flow: these solids are **insulators**.



The power is carried by aluminum wires, suspended by glass, polymer, or porcelain insulators.

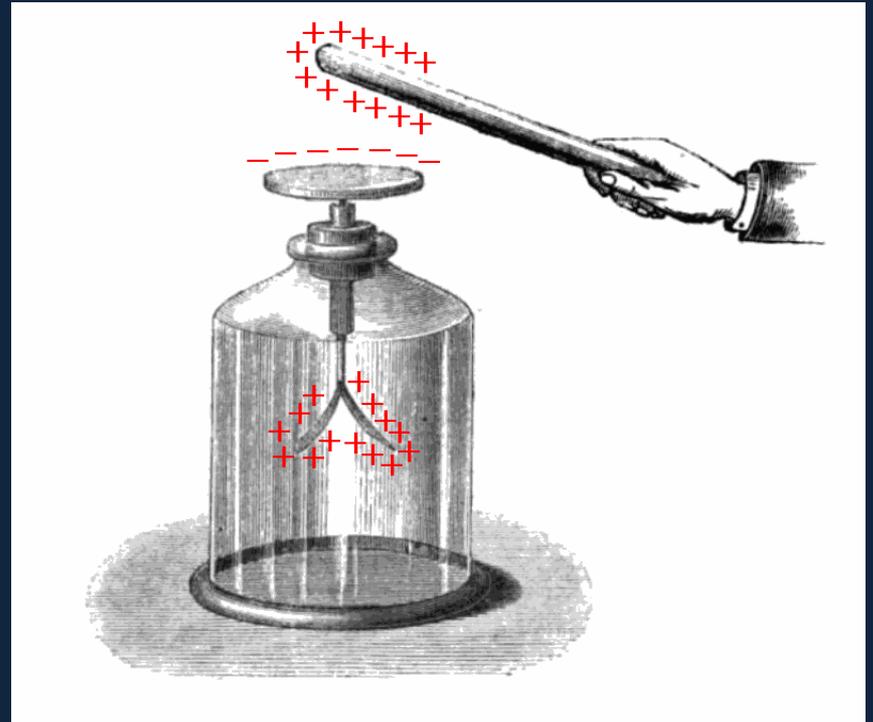
Attracting Neutrals

- A **neutral** conducting ball **will be attracted** by an electric charge: electrons will move around in the conductor to be closer to an external positive charge, so experiencing stronger force.
- For **insulators** there is sometimes a similar (weaker) effect: the molecules distort to put electrons closer.



The Electroscope

- Charge detector invented by an English clergyman in 1787. Two very thin strips of gold leaf hang side by side from a conducting rod.
- If a **+** charge is brought near, electrons move up the rod, leaving the two strips **positively charged**, so they repel each other.



Charging the Electroscope...

- By **conduction**: touch the top conductor with a positively charged object—this will leave it positively charged (electron deficient).
- By **induction**: while holding a positively charged object near, but not in contact, with the top, you touch the electroscope: negative charge will flow from the ground, through you, to the electroscope.

Clicker Question

- I charge the electroscope by **conduction**, **touching it with a charged object**, which I then remove. The leaves are apart.
- I now bring the charged object back close to the electroscope, but **not touching it**.
- The leaves:
 - A. Move further apart
 - B. Are not affected, if no touching
 - C. Move closer together.